

# IMAT

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International Medical Admissions Test③

FOR DEGREE COURSE IN INTERNATIONAL  
MEDICINE&SURGERY

The total number of questions is 60 and the duration is 100 minutes.

The scoring will be grades out of 90 Points.

The correct answers will be multiplied by +1.5 and the wrong ones by -0.4.

**2025-26 Academic year**

Reading skills and knowledge acquired during studies

**1. Which of the following authors wrote a work that primarily explores the biblical story of the Fall of Man, featuring Satan as a tragic figure?**

- A) Dante Alighieri – The Divine Comedy
- B) Geoffrey Chaucer – The Canterbury Tales
- C) William Blake – Songs of Innocence and of Experience
- D) John Milton – Paradise Lost
- E) Jonathan Swift – Gulliver's Travels

**2. Throughout the Middle Ages, the concept of law was heavily influenced by religious authority. Canon law, developed and administered by the Catholic Church, governed not only spiritual affairs but also significant aspects of public and private life. Unlike modern legal codes, medieval law was not always written or systematically codified. Instead, it was shaped by custom, papal decrees, and theological interpretations, and religious doctrine and moral teaching often took precedence over legal consistency.**

**Which of the following statements CANNOT be inferred from the text?**

- A) Religious beliefs had a strong influence on medieval legal systems.
- B) Moral values were often prioritized over legal consistency in medieval law.
- C) Laws in the Middle Ages were always formally written.
- D) The Catholic Church played a central role in shaping legal standards in the Middle Ages.
- E) Papal authority influenced both public and private life in medieval society.

**3. Which of the following countries is not a member of the Schengen Area?**

- A) Norway
- B) Switzerland
- C) Poland
- D) Ireland
- E) Austria

**4. In which of the following sentences is the underlined word not used as a relative pronoun or relative adverb?**

- A) I didn't understand the reason why she was crying.
- B) He said that he would come to the meeting.
- C) This is the hotel where we stayed last summer.
- D) The girl who is standing over there is my classmate.
- E) The car which broke down was brand new.

Logical reasoning and problem-solving

**5. Read the following information carefully:**

- All applicants who are shortlisted for an interview have submitted a completed application form.
- Some of those who submitted a completed application form did not meet the minimum criteria.
- No applicant who failed to meet the minimum criteria is shortlisted.

**Which of the following must be true?**

- A) All applicants who meet the minimum criteria are shortlisted
- B) Only applicants who are shortlisted submitted an application form
- C) No one is shortlisted without submitting a completed application form
- D) All applicants who submitted a form are shortlisted
- E) Some shortlisted applicants failed to meet the minimum criteria

**6. At a university science fair, 800 students participated in a survey about their subject preferences. The organizers noted that 60% of the students said they were currently taking biology, while 45% reported studying chemistry. Interestingly, 25% of all students said they were taking both subjects. During the closing ceremony, the organizers planned to distribute two types of subject-specific souvenirs. One for students who took only biology or only chemistry, and a different one for those who took both. Students who didn't take either subject received nothing.**

**Based on the survey data, how many students should receive the souvenir meant for those who are studying exactly one of the two subjects?**

- A) 320
- B) 400
- C) 440
- D) 500
- E) 600

**7. A cinema offers two types of tickets: adult tickets at €12 and student tickets at €8. On a particular day, the cinema sold a total of 120 tickets, resulting in €1,200 in revenue. Due to a system error, the exact number of each ticket type sold is unknown. However, it's known that the number of adult tickets sold was no more than twice the number of student tickets sold.**

**Based on this information, how many student tickets were sold?**

- A) 30
- B) 40
- C) 50
- D) 60
- E) 70

**8. Read the passage below and answer the question that follows:**

**"It is often assumed that if a product is labelled as 'natural', it must be healthy or safe. However, this assumption is flawed. Arsenic and cyanide are naturally occurring substances, yet both are highly toxic. Similarly, not all synthetic products are harmful — many life-saving medicines are produced artificially. The term 'natural' may appeal to emotions, but it says little about the actual safety or effectiveness of a product."**

**Which of the following is most strongly supported by the passage?**

- A) Products labelled as 'natural' are always less effective than synthetic ones
- B) The distinction between 'natural' and 'synthetic' is irrelevant in all contexts
- C) Emotional appeal is a reliable indicator of a product's value
- D) The label 'natural' does not guarantee a product's safety
- E) Synthetic substances should always be preferred over natural ones

**9. Look at the sequence of numbers below:**

2, 6, 12, 20, 30, ?

**Which number comes next in the sequence?**

- A) 36
- B) 42
- C) 44
- D) 48
- E) 52

Biology

**10. A plant cell is placed in a hypertonic solution. Over time, changes in the cell's water content and structure are observed.**

**Which of the following best describes what occurs in the plant cell?**

- A) The cell becomes turgid due to water intake.
- B) The cell wall dissolves and the cell bursts.
- C) The cell membrane pulls away from the cell wall due to water loss.
- D) The vacuole expands and stores more water.
- E) There is no net movement of water in or out of the cell.

**11. Which of the following statements is CORRECT?**

- A) Each antibody binds specifically to one type of antigen.
- B) Antibodies directly digest and eliminate viruses.
- C) Antigens are proteins made by lymphocytes to destroy pathogens.
- D) B cells produce antigens after infection.
- E) Antibodies are synthesized by red blood cells.

**12. During cellular respiration, the electron transport chain (ETC) creates a proton gradient across the inner mitochondrial membrane. What is the primary purpose of this gradient?**

- A) To produce NADH and FADH<sub>2</sub> for glycolysis
- B) To transport oxygen into the mitochondrial matrix
- C) To drive the synthesis of ATP via chemiosmosis
- D) To activate enzymes in the citric acid cycle
- E) To initiate the breakdown of fatty acids

**13. Which of the following sequences correctly describes the events that occur during skeletal muscle contraction?**

1. Calcium binds to troponin, exposing binding sites on actin.
2. The muscle fiber contracts.
3. Myosin heads bind to actin and perform power strokes using ATP.
4. A motor neuron releases acetylcholine at the neuromuscular junction.
5. The muscle cell membrane depolarizes, leading to calcium ion release from the sarcoplasmic reticulum.

**Select the correct order of these events.**

- A) 4 → 5 → 1 → 3 → 2
- B) 4 → 1 → 5 → 3 → 2
- C) 5 → 4 → 1 → 2 → 3
- D) 1 → 4 → 3 → 5 → 2
- E) 3 → 4 → 5 → 1 → 2

**14. The following data shows the concentration of a toxic chemical (ppm) across a simple food chain:**

Trophic level	Chemical concentration (ppm)
Phytoplankton	0.05
Zooplankton	0.1
Small fish	0.5
Large fish	1.8
Eagle	4.2

**Which of the following best illustrates this trend?**

- A) Increase in energy at higher trophic levels
- B) Increase in biomass through photosynthesis
- C) Accumulation of toxins in organisms higher in the food chain
- D) Decrease in genetic mutations
- E) Rapid extinction of predators

**15. A student investigates how different temperatures affect the activity of a plant enzyme. The rate of reaction increases with temperature up to a point, then sharply declines.**

**Which of the following statements are consistent with these observations?**

- 1. Enzyme activity increases as temperature enhances molecular motion
- 2. Extremely high temperatures denature the enzyme
- 3. Enzymes are consumed during the reaction, leading to decreased activity

- A) 1 and 2 only
- B) 1 only
- C) 2 and 3 only
- D) 3 only
- E) 1, 2 and 3

**16. Which of the following processes occur in the nucleus of eukaryotic cells?**

- 1. Transcription of DNA into pre-mRNA
- 2. Splicing of introns from RNA transcripts
- 3. Addition of 5' cap and poly-A tail to mRNA

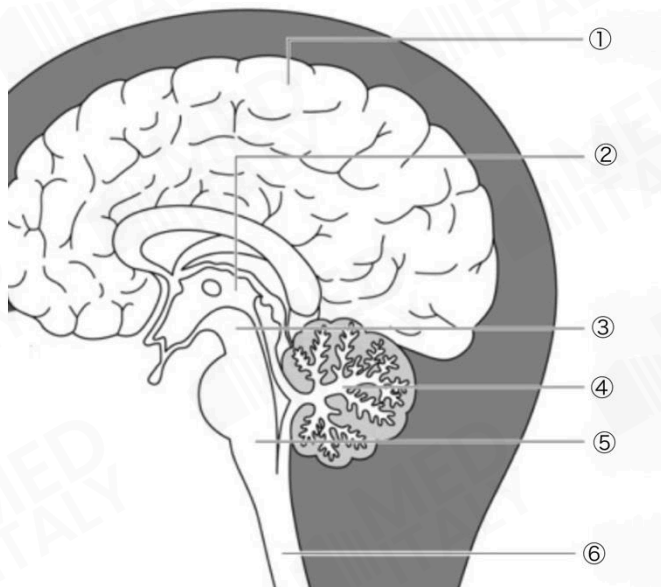
- A) 1 and 2 only
- B) 1 only
- C) 2 and 3 only
- D) 1, 2 and 3
- E) 3 only

**17. Which of the following statements about the plasma membrane is/are correct?**

1. Cholesterol modulates membrane fluidity by preventing tight packing of phospholipids at low temperatures.
2. Integral membrane proteins can move laterally within the lipid bilayer unless anchored to the cytoskeleton.
3. The plasma membrane is symmetric, with identical lipid and protein composition on both sides.
4. Small nonpolar molecules can diffuse through the lipid bilayer without the aid of transport proteins.

- A) 1 and 2 only  
 B) 1, 2, and 4 only  
 C) 2 and 3 only  
 D) 1, 2, 3, and 4  
 E) 1 and 3 only

**18.**



**Which of the following correctly matches brain regions to their typical functions?**

- A) ① = coordination of movement, ③ = visual reflexes, ④ = language processing  
 B) ① = memory and learning, ② = sensory relay station, ④ = balance coordination  
 C) ① = voluntary movement, ③ = auditory reflexes, ④ = olfaction  
 D) ① = respiration control, ③ = hunger regulation, ④ = visual reflex  
 E) ① = spinal reflexes, ③ = balance coordination, ④ = breathing regulation

**19. A woman is carrier by a rare autosomal recessive disease. It is also known that her partner has no alleles for the recessive disease.**

**What is the probability that their child will be a carrier?**

- A) 0%
- B) 25%
- C) 33%
- D) 50%
- E) 100%

**20. Which structure is not enclosed by a membrane?**

- A) Lysosome
- B) Nucleolus
- C) Golgi apparatus
- D) Peroxisome
- E) Endosome

**21. The endosymbiotic theory proposes that certain organelles, such as mitochondria, originated from free-living prokaryotes that were engulfed by ancestral eukaryotic cells and gradually became permanent, functional parts of the host.**

**To explore this idea, a biology teacher asked her students to design an experiment to test the theory. One student suggested comparing the DNA and ribosomes of mitochondria to those of bacteria.**

**Which of the following results would most strongly support the endosymbiotic origin of mitochondria?**

- A) Mitochondrial DNA contains introns, unlike bacterial DNA
- B) Mitochondrial DNA is double-stranded, while bacterial DNA is single-stranded
- C) Mitochondrial ribosomes are more similar to bacterial ribosomes than to those in the eukaryotic cytoplasm
- D) Mitochondria require nuclear genes for replication
- E) Mitochondria can survive independently when removed from a eukaryotic cell

**22. Glucose enters most cells via:**

- A) Active transport
- B) Secondary active transport
- C) Facilitated diffusion
- D) Endocytosis
- E) Simple diffusion



**23. Which molecule crosses the lipid bilayer most easily?**

- A) Glucose
- B)  $\text{Na}^+$
- C) Oxygen
- D)  $\text{Cl}^-$
- E) ATP

**24. Which molecule is the final electron acceptor in the light-dependent reactions?**

- A)  $\text{NADP}^+$
- B)  $\text{O}_2$
- C) ATP
- D)  $\text{CO}_2$
- E)  $\text{H}_2\text{O}$

**25. A group of students designed an experiment using radioactive isotopes to determine the source of oxygen released during photosynthesis. They used water labeled with  $^{18}\text{O}$  and carbon dioxide with  $^{16}\text{O}$ . After exposing the plant to light, the oxygen released was analyzed and found to be  $^{18}\text{O}$ .**

**What conclusion can be drawn from this result?**

- A) Oxygen released in photosynthesis originates from carbon dioxide
- B) Oxygen comes equally from water and carbon dioxide
- C) Oxygen originates from water molecules split during light-dependent reactions
- D) The plant absorbs oxygen from the air during photosynthesis
- E) Oxygen is a byproduct of glucose breakdown

**26. Under anaerobic conditions in muscle cells, glycolysis continues but oxidative phosphorylation does not.**

**Which of the following explains how glycolysis can proceed despite the absence of oxygen?**

- A)  $\text{NAD}^+$  is regenerated by converting pyruvate into lactic acid
- B) Pyruvate enters the Krebs cycle and is fully oxidized
- C) ATP is produced directly by the electron transport chain
- D) The cell begins to use stored glycogen to produce oxygen
- E) Glycolysis does not require any form of  $\text{NAD}^+$

**27. Why is oxygen essential in the ETC?**

- A) It phosphorylates ADP to form ATP
- B) It donates electrons to Complex I
- C) It maintains mitochondrial membrane potential
- D) It acts as the final electron acceptor
- E) It activates NADH dehydrogenase

**28. During transcription in eukaryotic cells, RNA polymerase reads the DNA template strand:**

- A) From 5' to 3' and synthesizes RNA in the 3' to 5' direction
- B) From 3' to 5' and synthesizes RNA in the 5' to 3' direction
- C) From 3' to 5' and synthesizes RNA in the 3' to 5' direction
- D) From 5' to 3' and synthesizes RNA in the 5' to 3' direction
- E) In both directions simultaneously

**29. Which of the following events occurs only in meiosis and not in mitosis?**

- A) Chromosome alignment at the metaphase plate
- B) Separation of sister chromatids
- C) DNA replication
- D) Pairing of homologous chromosomes
- E) Cytokinesis

**30. A mutation in a gene causes a codon to change from UAU (tyrosine) to UAA. What is the most likely result?**

- A) A silent mutation
- B) A frameshift mutation
- C) An extended polypeptide
- D) A truncated protein due to premature stop codon
- E) An amino acid substitution

**31. After completion of meiosis I in humans, each daughter cell is:**

- A) Diploid with duplicated chromosomes
- B) Diploid with single chromosomes
- C) Haploid with duplicated chromosomes
- D) Haploid with single chromosomes
- E) Aneuploid

**32. A red-flowered plant (RR) is crossed with a white-flowered plant (WW). All F<sub>1</sub> offspring are pink. This inheritance pattern is best described as:**

- A) Codominance
- B) Complete dominance
- C) Incomplete dominance
- D) Polygenic inheritance
- E) Epistasis

Chemistry

**33. Which of the following substances has the highest boiling point?**

- A) CH<sub>4</sub> (methane)
- B) CO<sub>2</sub> (carbon dioxide)
- C) NH<sub>3</sub> (ammonia)
- D) O<sub>2</sub> (oxygen)
- E) HCl (hydrogen chloride)

**34. Which of the following substances is a typical amphoteric oxide?**

- A) Na<sub>2</sub>O
- B) CO<sub>2</sub>
- C) Al<sub>2</sub>O<sub>3</sub>
- D) SO<sub>3</sub>
- E) CaO

**35. Which of the following reactions qualifies as a Lewis acid-base reaction, but cannot be classified as a Brønsted–Lowry acid-base reaction?**

- A)  $\text{Ag}^+ + 2\text{NH}_3 \rightarrow [\text{Ag}(\text{NH}_3)_2]^+$
- B)  $\text{CH}_3\text{COOH} + \text{NH}_3 \rightarrow \text{CH}_3\text{COO}^- + \text{NH}_4^+$
- C)  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
- D)  $\text{CO}_2 + \text{OH}^- \rightarrow \text{HCO}_3^-$
- E)  $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

**36. Which of the following statements about ionization energy are correct?**

- 1. Ionization energy increases across a period from left to right.
- 2. Ionization energy decreases down a group.
- 3. Noble gases have the lowest ionization energies in their periods.
- 4. The second ionization energy is always higher than the first.

- A) 1 and 2 only
- B) 2 and 3 only
- C) 1, 2 and 4 only
- D) 1, 3 and 4 only
- E) 1, 2, 3 and 4

**37. A student dilutes 100.0 mL of 1.0 M HCl to a final volume of 500.0 mL. What is the final concentration of HCl?**

- A) 0.10 M
- B) 0.20 M
- C) 0.25 M
- D) 0.50 M
- E) 1.0 M

**38. Which of the elements labeled a to e in the following periodic table form a strong base when combined with hydroxide ions?**

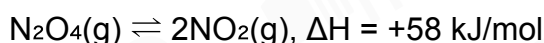
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- A) a,b and e  
B) a  
C) a and e  
D) c and e  
E) All of them

**39. Which of the following statements about enthalpy change ( $\Delta H$ ) is correct?**

- A) Endothermic reactions release heat to the surroundings  
B) Exothermic reactions have positive  $\Delta H$   
C) Breaking bonds is an exothermic process  
D) The overall enthalpy change is the sum of bond energies of bonds broken minus those formed  
E) Bond formation requires energy input, increasing  $\Delta H$

**40. For the equilibrium:**



**Which of the following changes would increase the yield of  $\text{NO}_2$ ?**

1. Increasing the temperature
2. Increasing the pressure
3. Removing  $\text{NO}_2$  as it forms
4. Adding an inert gas at constant pressure

- A) 1 and 3 only  
B) 1 and 2 only  
C) 2 and 3 only  
D) 3 and 4 only  
E) 1, 3 and 4 only

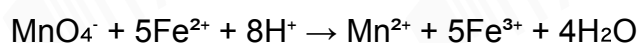
**41. The table below shows information for several molecules:**

Molecule	Electron pairs	Bonding pairs	Shape
NH <sub>3</sub>	4	3	a
CH <sub>4</sub>	4	4	b
H <sub>2</sub> O	4	2	c

**Which is the correct combination of shapes for a, b, and c?**

- A) Trigonal pyramidal, Tetrahedral, Bent
- B) Tetrahedral, Trigonal planar, Linear
- C) Linear, Tetrahedral, Trigonal pyramidal
- D) Trigonal planar, Trigonal planar, Bent
- E) Trigonal pyramidal, Trigonal planar, Linear

**42. A solution containing Fe<sup>2+</sup> is titrated with KMnO<sub>4</sub> under acidic conditions:**



**If 25.0 mL of 0.0200 mol/L KMnO<sub>4</sub> is required to react, how many mg of Fe<sup>2+</sup> were present? (M(Fe) = 55.8 g/mol)**

- A) 139.5 mg
- B) 55.8 mg
- C) 111.6 mg
- D) 27.9 mg
- E) 167.4 mg

**43. The formula of a molecule is CH<sub>3</sub>CH(OH)CH=CH<sub>2</sub>.**

**Which functional groups given below are present in the molecule?**

- 1. alkene
- 2. alcohol
- 3. aldehyde
- 4. ketone

- A) 1 and 2 only
- B) 1 and 3 only
- C) 2 and 4 only
- D) 3 and 4 only
- E) 1, 2 and 4 only

**44. Each of the following is a neutral polyatomic molecule. Determine the total number of electrons contained in the entire molecule, and select the one with the greatest number of electrons.**

- A)  $\text{C}_2\text{H}_5\text{NO}_2$  (glycine)
- B)  $\text{CH}_3\text{COOH}$  (acetic acid)
- C)  $\text{C}_2\text{H}_6\text{O}$  (ethanol)
- D)  $\text{H}_2\text{SO}_4$  (sulfuric acid)
- E)  $\text{C}_6\text{H}_6$  (benzene)

**45. A 10.0 mL of 0.10 mol/L hydrochloric acid solution is mixed with a 10.0 mL of 0.30 mol/L sodium hydroxide solution. Calculate the pH of the resulting mixture. Assume that the degree of dissociation of the acid and base is 100%, and the ionic product of water,  $K_w$ , is  $1.0 \times 10^{-14} \text{ (mol/L)}^2$ .**

- A) 10
- B) 11
- C) 12
- D) 13
- E) 14

**46. For each of the following compounds or ions, determine the oxidation number of the underlined atom and select the one with the highest oxidation number.**

- A)  $\text{H}\underline{\text{C}}\text{I}$
- B)  $\text{H}\underline{\text{C}}\text{I}\text{O}_4$
- C)  $\text{K}_2\underline{\text{C}}\text{r}_2\text{O}_7$
- D)  $\underline{\text{N}}\text{H}_4^+$
- E)  $\text{Ca}\underline{\text{S}}\text{O}_4$

**47. Assume that the cryoscopic constant (freezing point depression constant) of water is  $1.85 \text{ K}\cdot\text{kg/mol}$ .**

**When 2.4 g of urea  $[\text{CO}(\text{NH}_2)_2]$  is dissolved in 100 g of water, what is the freezing point of the resulting aqueous solution? ( $M(\text{CO}(\text{NH}_2)_2) = 60 \text{ g/mol}$ )**

- A)  $-0.40^\circ\text{C}$
- B)  $-0.55^\circ\text{C}$
- C)  $-0.74^\circ\text{C}$
- D)  $-1.11^\circ\text{C}$
- E)  $-1.48^\circ\text{C}$

Physics and Mathematics

**48. What is the minimum distance of the ellipse given by the equation:**

$$x^2 - 6x + 4y^2 + 8y + 9 = 0$$

**from the coordinate axes?**

- A) 1
- B) 2
- C) 3
- D) 4
- E) 0

**49. A committee of 3 members is to be formed from a group of 5 men and 4 women. What is the probability that the committee has at least one woman?**

- A)  $\frac{5}{7}$
- B)  $\frac{10}{21}$
- C)  $\frac{37}{42}$
- D)  $\frac{2}{3}$
- E)  $\frac{3}{5}$

**50. The expression  $\left(\frac{2x^3}{y^2}\right)^2 \times \frac{y^4}{4x^2}$  is equivalent to:**

- A)  $x^4 y^2$
- B)  $x^4$
- C)  $x^2$
- D)  $x^2 y^2$
- E)  $x^4 y^4$

**51. How many different integers,  $n$ , satisfy the inequality:**

$$|1.5n - 4| < 3$$

- A) 2
- B) 3
- C) 4
- D) 5
- E) 6

**52. Simplify:**  $\frac{\tan x + \frac{\cos x}{\sin x}}{\frac{1}{\cos x} + \frac{1}{\sin x}}$

- A)  $\sin x + \cos x$
- B)  $\frac{1}{\sin x \cos x}$
- C)  $\frac{1}{\sin x + \cos x}$
- D)  $\sin x \times \cos x$
- E)  $\tan x$

**53. Simplify:**  $\ln(x^2/4y) + \ln(xy) + \ln 8$

- A)  $3\ln x + 2\ln 2$
- B)  $\ln((x^2 + 4xy^2 + 32y)/4y)$
- C)  $3\ln x + 2\ln y + \ln 32$
- D)  $3\ln x + \ln 2$
- E)  $4\ln x + \ln 2$

**54. Simplify:**  $\sqrt[3]{\frac{x^3 + 6x^2 + 12x + 8}{x^3}}$

- A)  $\frac{x+2}{x}$
- B)  $\frac{2+3x}{x^2}$
- C)  $\frac{(x+2)^3}{x^3}$
- D)  $\frac{(x+4)^3}{x^3}$
- E) Cannot be simplified without factoring numerator

**55. Which of the following statements about uniform circular motion is TRUE?**

- A) The net force is always tangential to the motion
- B) The tension in the string is directed away from the center
- C) The angular velocity changes with time
- D) The object moves with constant speed but changing velocity
- E) The acceleration vector is zero



**56. Two objects with masses  $m_1$  and  $m_2$  are separated by a distance  $r$  in space, free from any other forces. According to Newton's law of universal gravitation, which of the following statements is false?**

- A) The gravitational force between the objects is proportional to  $\frac{1}{r^2}$
- B) If both masses are doubled, the gravitational force becomes four times stronger
- C) If the distance  $r$  is doubled, the gravitational force becomes one-fourth as strong
- D) The gravitational force acts equally on both masses but in opposite directions
- E) The gravitational force depends on the relative velocity of the two objects

**57. A cylindrical conductor has length  $L=1.0$  m cross-sectional area  $A=5.0 \times 10^{-7} \text{ m}^2$ , and a current of 2.0 A flows through it under a 12 V potential. What is its resistance?**

- A)  $0.5 \Omega$
- B)  $1.0 \Omega$
- C)  $2.0 \Omega$
- D)  $6.0 \Omega$
- E)  $12.0 \Omega$

**58. An airship is flying horizontally at a constant speed of 20 m/s. A small ball is released quietly from the airship, and it hits a target on the ground exactly 5.0 seconds later. Assume gravitational acceleration is  $9.8 \text{ m/s}^2$ . How far horizontally from the target was the ball released?**

- A) 50 m
- B) 75 m
- C) 100 m
- D) 120 m
- E) 150 m

**59. When 1.0 kg of heavy oil is burned, it produces  $4.2 \times 10^7 \text{ J}$  of heat. A diesel engine with a thermal efficiency of 40% consumes 20 kg of heavy oil. How much work (in joules) is done by the engine?**

- A)  $2.8 \times 10^8 \text{ J}$
- B)  $3.3 \times 10^8 \text{ J}$
- C)  $4.2 \times 10^8 \text{ J}$
- D)  $5.6 \times 10^8 \text{ J}$
- E)  $6.7 \times 10^8 \text{ J}$

60. A long straight wire carries a current of 15.7 A. At a point 20 cm away from the wire, there is a circular coil of radius 10 cm centered at that point. The circular coil has 5 turns and lies in the same plane as the straight wire.

Assume the magnetic permeability of air is  $\mu_0 = 1.3 \times 10^{-6} \text{ N/A}^2$

In order to make the net magnetic field at the center of the circular coil zero, what magnitude of current should be passed through the circular coil?

- A) 0.25 A
- B) 0.50 A
- C) 0.75 A
- D) 1.00 A
- E) 1.25 A

—THE END—